

**CORROSION AND GAS HYDRATE INHIBITORS HAVING
IMPROVED WATER SOLUBILITY AND
INCREASED BIODEGRADABILITY**

- [0001] The present invention relates to an additive and to a process for corrosion inhibition and gas hydrate inhibition on and in equipment for extracting and transporting hydrocarbons in crude oil extraction and processing.
- [0002] In industrial processes in which metals come into contact with water or else with biphasic oil-water systems, there is the risk of corrosion. This is particularly marked when the aqueous phase, as in the case of crude oil production and processing operations, has a high salt content or, as a result of dissolved acidic gases such as carbon dioxide or hydrogen sulfide, is acidic. It is therefore not possible to exploit a deposit and to process crude oil without special additives to protect the equipment used.
- [0003] Although suitable corrosion protectors for crude oil extraction and processing have been known for some time, they will become unacceptable for offshore applications in the future for reasons of environmental protection.
- [0004] As typical prior art corrosion inhibitors, amides, amido amines and imidazolines of fatty acids and polyamines have extremely good oil solubility and are therefore only present in the corrosive aqueous phase in low concentration as a consequence of poor partitioning equilibria. Accordingly, these products have to be used in high dosage despite their poor biodegradability.
- [0005] DE-A-199 30 683 describes corresponding amido amines/imidazolines which are obtained by reacting alkylpolyglycol ether carboxylic acids

with polyamines and which, as a consequence of better partitioning, can be used in low concentrations.

- [0006] Quaternary alkylammonium compounds (quats) constitute alternative prior art corrosion protectors which have not only corrosion-inhibiting but also biostatic properties. Despite improved water solubility, the quats, in comparison to the imidazolines for example, exhibit distinctly reduced film persistence and therefore likewise lead to effective corrosion protection only in high dosage. The strong algae toxicity and the moderate biodegradability restrict the use of quats ever more to ecologically insensitive fields of application, for example onshore.
- [0007] EP-B-0 946 788 describes a process for protecting metal surfaces against corrosion using ester quats which, it is disclosed, have good biodegradability and low aquatic toxicity.
- [0008] EP-A-0 320 769 discloses optionally quaternized fatty acid esters of oxyalkylated alkylamino alkylenamines and their use as corrosion inhibitors.
- [0009] EP-B-0 212 265 describes quaternary polycondensates of alkoxyated alkylamines and dicarboxylic acids and their use as corrosion inhibitors and demulsifiers in crude oils.
- [00010] EP-B-0 446 616 describes ampholytes and betaines based on fatty acid amido alkylamines which have very good corrosion protection and significantly reduced algae toxicity under the given test conditions.
- [00011] Gas hydrates are crystalline inclusion compounds of gas molecules in water which form under certain temperature and pressure conditions (low temperature and high pressure). The water molecules form cage structures around the appropriate gas molecules. The lattice structure

formed from the water molecules is thermodynamically unstable and is always stabilized by the incorporation of gas molecules. Depending on pressure and gas composition, these icelike compounds can exist even to above the freezing point of water (up to above 25°C).

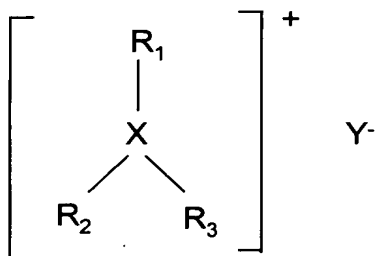
- [00012] In the crude oil and natural gas industry, great significance attaches in particular to the gas hydrates which form from water and the natural gas constituents methane, ethane, propane, isobutane, n-butane, nitrogen, carbon dioxide and hydrogen sulfide. Especially in modern natural gas extraction, the existence of these gas hydrates constitutes a great problem, especially when wet gas or multiphasic mixtures of water, gas and alkane mixtures are subjected to low temperatures under high pressure. As a consequence of their insolubility and crystalline structure, the formation of gas hydrates leads here to the blockage of a wide variety of extraction equipment such as pipelines, valves or production equipment in which wet gas or multiphasic mixtures are transported over long distances, as occurs especially in colder regions of the earth or on the seabed.
- [00013] In addition, gas hydrate formation can also lead to problems in the course of the drilling operation to develop new gas or crude oil deposits at the appropriate pressure and temperature conditions by the formation of gas hydrates in the drilling fluids.
- [00014] In order to prevent such problems, gas hydrate formation in gas pipelines, in the course of transport of multiphasic mixtures or in drilling fluids, can be suppressed by using relatively large amounts (more than 10% by weight, based on the weight of the aqueous phase) of lower alcohols such as methanol, glycol or diethylene glycol. The addition of these additives has the effect that the thermodynamic limit of gas hydrate formation is shifted to lower temperatures and higher pressures (thermodynamic inhibition). However, the addition of these

thermodynamic inhibitors causes serious safety problems (flashpoint and toxicity of the alcohols), logistical problems (large storage tanks, recycling of these solvents) and accordingly high costs, especially in offshore extraction.

[00015] Attempts are therefore being made today to replace thermodynamic inhibitors by adding additives in amounts of < 2% in temperature and pressure ranges in which gas hydrates can form. These additives either delay gas hydrate formation (kinetic inhibitors) or keep the gas hydrate agglomerates small and therefore pumpable, so that they can be transported through the pipeline (agglomerate inhibitors or antiagglomerates). The inhibitors used either prevent nucleation and/or the growth of the gas hydrate particles, or modify the hydrate growth in such a way that relatively small hydrate particles result.

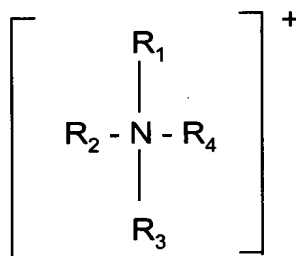
[00016] The gas hydrate inhibitors which have been described in the patent literature, in addition to the known thermodynamic inhibitors, are a multitude of monomeric and also polymeric substance classes which are kinetic inhibitors or agglomeration inhibitors. Of particular significance in this context are polymers having a carbon backbone which contain both cyclic (pyrrolidone or caprolactam radicals) and acyclic amide structures in the side groups.

[00017] EP-B-0 736 130 discloses a process for inhibiting gas hydrates which entails feeding a substance of the formula



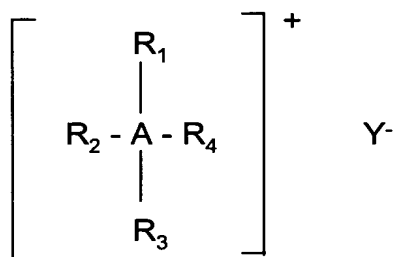
where $X = S, N-R_4$ or $P-R_4$, R_1, R_2 and $R_3 =$ alkyl having at least 4 carbon atoms, $R_4 = H$ or an organic radical, and $Y =$ anion.

[00018] This therefore includes compounds of the formula



where R_4 may be any desired radical, but the R_1 to R_3 radicals have to be alkyl radicals having at least 4 carbon atoms.

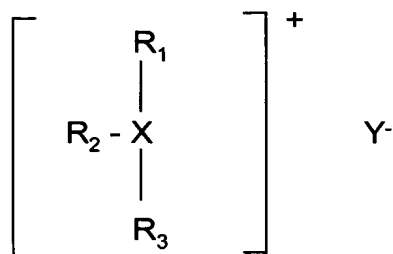
[00019] EP-B-0 824 631 discloses a process for inhibiting gas hydrates which entails feeding a substance of the formula



where $R_1, R_2 =$ linear/branched alkyl radicals having 4 or 5 carbon atoms, $R_3, R_4 =$ organic radicals having at least 8 carbon atoms and $A =$ nitrogen or phosphorus. Y^- is an anion. Two of the R_1 to R_4 radicals have to be linear or branched alkyl radicals having 4 or 5 carbon atoms.

[00020] US-5 648 575 discloses a process for inhibiting gas hydrates. The

process comprises the use of a compound of the formula



where R¹, R² are linear or branched alkyl groups having at least 4 carbon atoms, R³ is an organic radical having at least 4 atoms, X is sulfur, NR⁴ or PR⁴, R⁴ is hydrogen or an organic radical, and Y is an anion. The document discloses only those compounds which have at least two alkyl radicals having at least 4 carbon atoms.

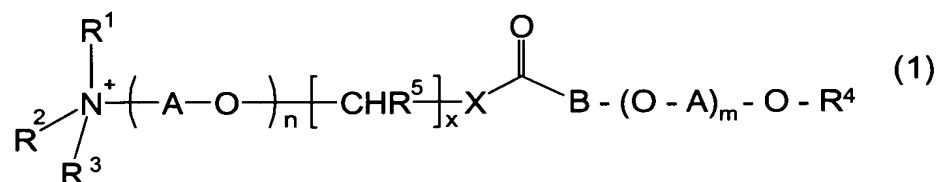
- [00021] US-6 025 302 discloses polyetheramine ammonium compounds as gas hydrate inhibitors whose ammonium nitrogen atom, in addition to the polyetheramine chain, bears 3 alkyl substituents.
- [00022] WO-99/13197 discloses ammonium compounds as gas hydrate inhibitors which have at least one alkoxy group esterified with alkylcarboxylic acids. The advantages of using ether carboxylic acids are not disclosed.
- [00023] WO-01/09082 discloses a process for preparing quaternary amines which, however, bear no alkoxy groups, and their use as gas hydrate inhibitors.
- [00024] WO-00/078 706 discloses quaternary ammonium compounds as gas hydrate inhibitors which, however, bear no carbonyl radicals.
- [00025] EP-B-0 914 407 discloses the use of trisubstituted amine oxides as gas

hydrate inhibitors.

- [00026] DE-C-10114638 and DE-C-19920152 describe the use of modified ether carboxamides as gas hydrate inhibitors. The use of quaternized ether carboxylic acid derivatives is not disclosed.
- [00027] It is an object of the present invention to find novel corrosion inhibitors which, coupled with equally good or improved corrosion protection, offer not only optimum water solubility, more rapid film formation and therefore improved film persistence, but also improved biodegradability in comparison to the prior art corrosion inhibitors.
- [00028] It is a further object of the present invention to find improved additives which not only slow the formation of gas hydrates (kinetic inhibitors) but also keep gas hydrate agglomerates small and pumpable (antiagglomerates), in order to thus ensure a broad spectrum of application with a high action potential. In addition, it should be possible to replace the thermodynamic inhibitors used currently (methanol and glycols) which cause considerable safety problems and logistical problems.
- [00029] Prior art gas hydrate inhibitors are commonly coadditized with corrosion inhibitors, in order to prevent corrosion of the transport and extraction equipment. As a consequence of the frequent lack of immediate compatibility of gas hydrate inhibitor and corrosion protector in the course of formulation, there is additional work for the user. It would be a significant advantage over the prior art if coadditization with corrosion inhibitors were no longer obligatory.
- [00030] It has now been found that, surprisingly, quaternary alkylaminoalkyl/alkoxy esters and quaternary alkylaminoalkyl/alkoxy amides of ether carboxylic acids exhibit excellent action as corrosion

inhibitors and gas hydrate inhibitors, and also improved film persistence and good biodegradability.

[00031] The present invention therefore provides the use of compounds of the formula (1)



where

R^1, R^2 are each independently C_1 - to C_{22} -alkyl, C_2 - to C_{22} -alkenyl, C_6 - to C_{30} -aryl or C_7 - to C_{30} -alkylaryl,

R^3 is C_1 - to C_{22} -alkyl, C_2 - to C_{22} -alkenyl, C_6 - to C_{30} -aryl or C_7 - to C_{30} -alkylaryl, $-\text{CHR}^6\text{COO}^-$ or $-\text{O}^-$,

A is a C_2 - to C_4 -alkylene group,

B is a C_1 - to C_{10} -alkylene group,

X is O or NR^7

R^6, R^7 are each independently hydrogen, C_1 - to C_{22} -alkyl, C_2 - to C_{22} -alkenyl, C_6 - to C_{30} -aryl or C_7 - to C_{30} -alkylaryl,

R^4 is a C_1 - to C_{50} -alkyl, C_2 - to C_{50} -alkenyl radical, C_6 - to C_{50} -aryl or C_7 - to C_{50} -alkylaryl,

R^5 is hydrogen, $-\text{OH}$ or a C_1 - to C_4 -alkyl radical,

n, m are each independently a number from 0 to 30,

x is a number from 1 to 6,

as corrosion and gas hydrate inhibitors.

[00032] The invention further provides a method for inhibiting corrosion on metal surfaces, in particular ferrous surfaces, by adding at least one compound of the formula (1) to a corrosive system which is in contact

with the metal surfaces.

- [00033] The invention further provides a method for inhibiting gas hydrates by adding at least one compound of the formula (1) to a system of water and hydrocarbons which tends to the formation of gas hydrates.
- [00034] The invention further provides the compounds of the formula (1).
- [00035] For the purposes of this invention, corrosive systems are preferably liquid/liquid or liquid/gaseous multiphase systems consisting of water and hydrocarbons which comprise corrosive constituents, such as salts and acids, in free and/or dissolved form. The corrosive constituents may also be gaseous, for instance hydrogen sulfide and carbon dioxide.
- [00036] For the purposes of this invention, hydrocarbons are organic compounds which are constituents of the crude oil/natural gas, and their secondary products. For the purposes of this invention, hydrocarbons are also volatile hydrocarbons, for example methane, ethane, propane, butane. For the purposes of this invention, they also include the further gaseous constituents of crude oil/natural gas, for instance hydrogen sulfide and carbon dioxide.
- [00037] A may be straight-chain or branched and is preferably an ethylene or propylene group, in particular an ethylene group. The alkoxy groups denoted by $(A-O)_n$ and $(O-A)_m$ may also be mixed alkoxy groups.
- [00038] B may be straight-chain or branched and is preferably a C_2 - to C_4 -alkylene group, in particular an ethylene or propylene group.
- [00039] R^1 and R^2 are preferably each independently an alkyl or alkenyl group of from 2 to 14 carbon atoms, in particular those groups having from 3

to 8 carbon atoms and especially butyl groups.

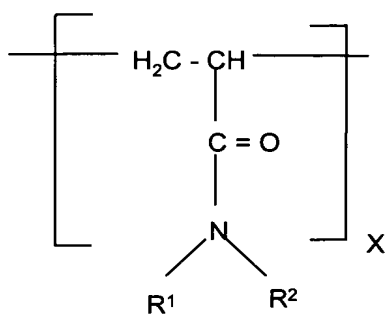
- [00040] R^3 is preferably an alkyl or alkenyl group having from 1 to 12 carbon atoms, in particular having from 1 to 4 carbon atoms.
- [00041] R^4 is preferably an alkyl or alkenyl group having from 4 to 30 carbon atoms, in particular having from 6 to 18 carbon atoms. The alkyl or alkenyl group may be straight-chain or branched, preferably branched.
- [00042] R^5 , R^6 and R^7 are preferably each independently hydrogen. In a further preferred embodiment, they are simultaneously each hydrogen.
- [00043] n is preferably a number in the range from 1 to 10, more preferably in the range from 2 to 6.
- [00044] m , independently of n , is a number in the range from 0 to 30, preferably in the range from 1 to 10, and more preferably in the range from 2 to 6. x is preferably the numbers 2 and 3.
- [00045] When R^3 is C_1 - to C_{22} -alkyl, C_2 - to C_{22} -alkenyl or C_6 - to C_{30} -alkylaryl, suitable counterions for the compounds of formula (1) are any anions which do not impair the solubility of the compounds of the formula (1) in the organic-aqueous mixed phases. Such counterions are, for example, methylsulfate ions (methosulfate) or halide ions.
- [00046] When R^3 is $-\text{CHR}^5\text{-COO}^-$ or $-\text{O}^-$, the compounds of the formula (1) are betaines and amine oxides respectively and, as internal salts (ampholytes), have an intramolecular counterion.
- [00047] The compounds according to the invention can be used alone or in combination with other known corrosion inhibitors and/or gas hydrate inhibitors. In general, sufficient corrosion inhibitor and/or gas hydrate

inhibitor according to the invention is used to obtain sufficient corrosion protection and protection from gas hydrate formation under the given conditions.

- [00048] Preferred use concentrations of the corrosion inhibitors based on the pure compounds of the invention are from 5 to 5 000 ppm, preferably from 10 to 1000 ppm, in particular from 15 to 150 ppm.
- [00049] The gas hydrate inhibitors are generally used in amounts between 0.01 and 5% by weight of the pure compounds according to the invention based on the aqueous phase, preferably between 0.05 and 2% by weight.
- [00050] Also particularly suitable as corrosion inhibitors and/or gas hydrate inhibitors are mixtures of the products according to the invention with other prior art corrosion inhibitors and/or gas hydrate inhibitors.
- [00051] Particularly suitable corrosion inhibitors and therefore a preferred embodiment of this invention are mixtures of the compounds of the formula (1), such as amido amines and/or imidazolines of fatty acids and polyamines and their salts, quaternary ammonium salts, ethoxylated/propoxylated amines, amphoglycinates and -propionates, betaines or compounds described in DE-A-19930683.
- [00052] Particularly suitable gas hydrate inhibitors and therefore a preferred embodiment of this invention are mixtures of the compounds of the formula (1) with one or more polymers having a carbon backbone obtained by polymerization and amide bonds in the side chains. These include in particular polymers such as polyvinylpyrrolidone, polyvinylcaprolactam, polyisopropylacrylamide, polyacryloylpyrrolidine, copolymers of vinylpyrrolidone and vinylcaprolactam, copolymers of vinylcaprolactam and N-methyl-N-vinylacetamide, and also terpolymers

of vinylpyrrolidone, vinylcaprolactam and further anionic, cationic and neutral comonomers having a vinylic double bond, such as 2-dimethylaminoethyl methacrylate, 1-olefins, N-alkylacrylamides, N-vinylacetamide, acrylamide, sodium 2-acrylamido-2-methyl-1-propanesulfonate (AMPS) or acrylic acid. Also suitable are mixtures with homo- and copolymers of N,N-dialkylacrylamides such as N-acryloylpyrrolidine, N-acryloylmorpholine and N-acryloylpiperidine, or N-alkylacrylamides such as isopropylacrylamide, and also both homo- and copolymers of alkoxyalkyl-substituted (meth)acrylic esters. Likewise suitable are mixtures with alkylpolyglycosides, hydroxyethylcellulose, carboxymethylcellulose and also other ionic or nonionic surfactant molecules.

[00053] In a preferred embodiment of the invention, the gas hydrate inhibitors according to the invention are used in a mixture with polymers which are disclosed in WO-A-96/08672. These polymers are those which have structural units of the formula

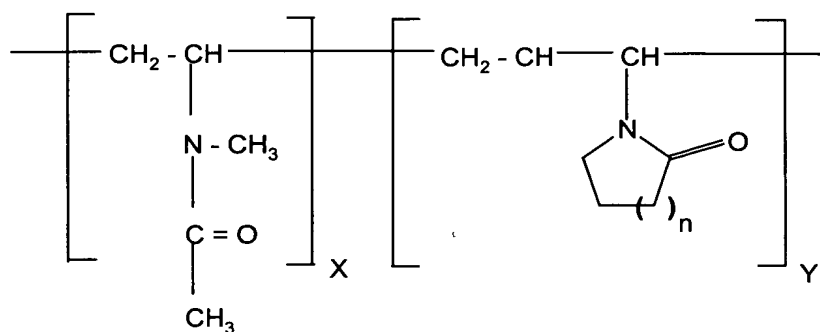


where R^1 is a hydrocarbon group having from 1 to 10 carbon atoms and from 0 to 4 heteroatoms selected from N, O and S, R^2 is a hydrocarbon group as defined for R^1 , and X is the average number of the repeating units, the latter being such that the polymer has a molecular weight of from 1 000 to 6 000 000. For use in the context of the present invention,

polyisopropylacrylamides and polyacryloylpyrrolidines are particularly suitable.

[00054] In a further preferred embodiment of the invention, the gas hydrate inhibitors according to the invention are used in a mixture with polymers which are disclosed in WO-A-96/41785.

[00055] This document discloses gas hydrate inhibitors including structural units



where n is a number from 1 to 3, and x and y are the number of repeating units which is such that the molecular weight of the polymer is between 1 000 and 6 000 000. For use in the context of the present invention, copolymers of N-vinylcaprolactam and N-vinyl-N-methylacetamide or vinylpyrrolidone are particularly suitable.

[00056] In a further preferred embodiment of the invention, the gas hydrate inhibitors according to the invention are used in a mixture with other polymers which are disclosed in WO-A-94/12761. The document discloses additives which are for preventing the formation of gas hydrates and contain polymers including cyclic constituents having from 5 to 7 ring members. For use in the context of the present invention, polyvinylcaprolactam, polyvinylpyrrolidone and hydroxyethylcellulose in particular are suitable.

- [00057] Also particularly suitable are mixtures of the polymers according to the invention with gas hydrate inhibitors based on maleic anhydride, as described in WO-A-97/6506, in particular with maleic anhydride copolymers reacted with mono- and/or diamines. Among these particular preference is given in particular to modified vinyl acetate-maleic anhydride copolymers.
- [00058] In a further preferred embodiment of the invention, the gas hydrate inhibitors according to the invention are used in a mixture with other polymers which are disclosed in DE-C-10059816. The document discloses additives which are for preventing the formation of gas hydrates and contain polymers including homo- and also copolymers which contain alkoxyalkyl-substituted (meth)acrylic esters.
- [00059] When mixtures are used, the concentration ratios between the gas hydrate inhibitors according to the invention and the mixed-in components are from 90:10 to 10:90 percent by weight, and preference is given to mixtures in the ratios from 75:25 to 25:75, and in particular from 60:40 to 40:60.
- [00060] Like their mixtures with other gas hydrate inhibitors, the compounds can be added to the multiphase mixture which is prone to hydrate formation in the course of crude oil and natural gas extraction or in the course of provision of drilling muds using common equipment such as injection pumps or the like; as a consequence of the good solubility of the polymers according to the invention, there is rapid and uniform distribution of the inhibitor in the aqueous phase or the condensate phase tending to hydrate formation.
- [00061] The compounds according to the invention can be prepared by reacting alkoxyated alkylamines, alkylaminoalkylenamines or else alkylaminoalkyletheramines with ether carboxylic acid derivatives to

give the corresponding ether carboxylic esters, or ether carboxamides. Subsequently, quaternization is effected using suitable alkylating agents.

- [00062] The preparation of alkoxyated alkylamines, alkylaminoalkylenamines or else alkylaminoalkyletheramines is described in the prior art.
- [00063] The alkoxyated alkylamines used are based on dialkylamines having C₁- to C₂₂-alkyl radicals or C₂- to C₂₂-alkenyl radicals, preferably C₃- to C₈-dialkylamines. Suitable dialkylamines are, for example, di-n-butylamine, diisobutylamine, dipentylamine, dihexylamine, dioctylamine, dicyclopentylamine, dicyclohexylamine, diphenylamine, dibenzylamine.
- [00064] The alkylamines are generally reacted with ethylene oxide, propylene oxide, butylene oxide or mixtures of different such alkylene oxides, although preference is given to ethylene oxide or mixtures of ethylene oxide and propylene oxide. Based on the alkylamines, 1-30 mol of alkylene oxide are added, preferably 1-10 mol, more preferably 2-6 mol.
- [00065] The alkoxylation proceeds without solvent, but can also be carried out in solution. Suitable solvents for the alkoxylation are inert ethers such as dioxane, tetrahydrofuran, glyme, diglyme and MPEGs.
- [00066] In general, the alkoxylation is carried out uncatalyzed in the first reaction step up to > 95% by weight of tertiary nitrogen. Higher alkoxylation is effected after addition of basic compounds as catalysts. Useful basic compounds are alkali metal/alkaline earth metal hydroxides or alkoxides (sodium methoxide, sodium ethoxide, potassium tert-butoxide), but preference is given to alkali metal hydroxides, particularly sodium hydroxide or potassium hydroxide.
- [00067] The alkylaminoalkylenamines used are based substantially on

dialkylaminoalkylenamines having C_1 - to C_{22} -alkyl radicals or C_2 - to C_{22} -alkenyl radicals, preferably tertiary C_1 - to C_8 -dialkylaminoalkylenamines. Particularly suitable are, for example, N,N-dibutylaminopropylamine, N,N-diethylaminopropylamine, N,N-dimethylaminopropylamine, N,N-dimethylaminobutylamine, N,N-dimethyl-aminoethylamine, N,N-dimethylaminodecylamine, N,N-dibutylaminoethylamine and N,N-dimethylamino-2-hydroxypropylamine.

[00068] The parent ether carboxylic acids can be prepared by known routes, either by reaction of the terminal $-CH_2OH$ function of a C_1 - C_{50} alcohol or alkylphenol reacted with alkylene oxide to give the carboxylic acid, by alkylation with chloroacetic acid derivatives by the Williamson ether synthesis, or by oxidation.

[00069] Suitable base alcohols are more preferably C_4 - C_{30} fatty alcohols, such as butanol, isobutanol, pentanol, hexanol, n-octanol, dodecanol, tetradecanol, hexadecanol, octadecanol and fatty alcohol cuts of natural or synthetic origin, such as oleyl alcohol, stearyl alcohol, behenyl alcohol and oxo alcohols of varying chain length. Likewise suitable are branched alcohols such as 2-ethylhexanol, isononyl alcohol, isodecanol, isotridecanol, isooctadecanol or Guerbet alcohols, or else alkylphenols having a C_1 - C_{12} -alkyl radical such as nonylphenol.

[00070] These alcohols are reacted with alkylene oxides, such as ethylene oxide, propylene oxide or butylene oxide, or mixtures thereof to give corresponding glycol ethers, and particular preference is given to ethylene oxide. The products are reacted preferably with 1-30 mol of alkylene oxide, preferably 1-10 mol of ethylene oxide.

[00071] The ether carboxylic acid is prepared from these alkylene glycol monoalkyl ethers by oxidation with atmospheric oxygen in the presence

of catalysts or by oxidation with hypochlorite or chlorite, with and without catalysts.

- [00072] However, particular preference is given to alkylating the glycol ethers with chloroacetic acid derivatives, preferably with sodium chloroacetate and sodium hydroxide solution by the Williamson method. The free carboxylic acid is obtained from the alkaline alkylation mixture by acidifying with mineral acid (hydrochloric, sulfuric acid) and heating above the cloud point and removing the organic phase. The ether carboxylic acids are generally reacted with the alkoxyated alkylamines, alkylaminoalkylenamines or alkylaminoalkyl ether amines at 60 - 240°C, preferably at 120 – 200°C, in such a way that, in some cases depending on the ether carboxylic acid derivative used, there is complete condensation to the corresponding ether carboxylic ester, or ether carboxamide, with elimination of water of reaction or of alcohol. The degree of reaction can be followed by determination of the acid number, hydrolysis number and/or by the determination of the base and/or amide nitrogen.
- [00073] The reaction proceeds without solvent, but can also preferably be carried out in solution. The use of solvents is required especially when high conversions and low acid numbers of the resulting reaction products are pursued. Suitable solvents for the preparation are organic compounds which azeotropically remove the water of reaction. In particular, aromatic solvents or solvent mixtures, or alcohols, can be used. Particular preference is given to 2-ethylhexanol. The reaction is then effected at the boiling point of the azeotrope.
- [00074] According to the prior art, the esterifications and amidations can be accelerated by addition of suitable acid catalysts having a pK_a of less than or equal to 5, for example mineral acids or sulfonic acids. Particular preference is given to alkylstannic acids.

- [00075] For the preparation of the compounds according to the invention, the alkylaminoalkyl/alkoxy esters or alkylaminoalkyl/alkoxy amides are appropriately quaternized in a subsequent reaction step. The quaternization may be effected by appropriate alkylating agents at from 50 to 150°C. Suitable alkylating agents are alkyl halides and alkyl sulfates, preferably methyl chloride, methyl iodide, butyl bromide and dimethyl sulfate.
- [00076] For the preparation of the betaines according to the invention, reaction is effected with a halocarboxylic acid and a base, preferably chloroacetic acid and sodium hydroxide. This may be effected by initially charging the alkylamino amides with from 50 to 125 mol% of halocarboxylic acid at 40°C and reacting at from 40 to 100°C by adding the base and the amount remaining up to 125 mol% of halocarboxylic acid, all at once or in portions.
- [00077] The basic compounds used may be alkali metal/alkaline earth metal hydroxides or alkoxides (sodium methoxide, sodium ethoxide, potassium tert-butoxide), but preferably alkali metal hydroxides, particularly sodium hydroxide or potassium hydroxide, in particular their aqueous solutions.
- [00078] The amine oxides according to the invention are prepared by existing prior art processes, preferably by oxidation of the appropriate tertiary amine group with peroxides or peracids, preferably with hydrogen peroxide.
- [00079] The reactions to give the quaternary alkylaminoalkyl/alkoxy esters or alkylaminoalkyl/alkoxy amides proceed without solvent, but can also preferably be carried out in solution. Suitable solvents for the preparation of quats, betaines and amine oxides are inert alcohols such as isopropanol, 2-ethylhexanol, or inert ethers such as tetrahydrofuran,

glyme, diglyme and MPEGs.

- [00080] Depending on the given requirements, the solvent used may remain in the product according to the invention or has to be removed distillatively.

Examples:

- [00081] General method for the preparation of N,N'-dialkylaminoalkyl ether carboxylates
- [00082] A stirred apparatus equipped with a reflux condenser was initially charged with 0.3 mol of the appropriate ether carboxylic acid (based on hydrolysis number) with nitrogen purging and heated to 60°C. 0.3 mol of the appropriate N,N-dialkylaminoethanol and 0.5% by weight of butylstannic acid were then added. The reaction mixture was heated to 120°C, in the course of which the first water of condensation passed over. The reaction temperature was then increased continuously to 180°C over 15 h until no more water of reaction condensed out and an acid number of < 10 mg KOH/g had been attained. Finally, the reaction product was clarified using a pressure filter press with Seitz T 2600.
- [00083] Example 1 (N,N-dibutylamino-N-ethyl oleyl+5 EO-ether carboxylate)
203.9 g of oleyl+5 EO-ether carboxylic acid (HN = 82.6 mg KOH/g) and 52.0 g of dibutylaminoethanol (DBAE, OHN = 323.8 mg KOH/g) were used to obtain approx. 230 g of N,N-dibutylamino-N-ethyl oleyl+5 EO-ether carboxylate having AN = 3.5 mg KOH/g, HN = 75.3 mg KOH/g and basic N = 1.64%.
- [00084] Example 2 (N,N-dibutylamino-N-ethyl oleyl+8 EO-ether carboxylate)
235.9 g of oleyl+8 EO-ether carboxylic acid (HN = 71.4 mg KOH/g) and 52.0 g of dibutylaminoethanol (DBAE, OHN = 323.8 mg KOH/g) were

used to obtain approx. 235 g of N,N-dibutylamino-N-ethyl oleyl+8 EO-ether carboxylate having AN = 4.3 mg KOH/g, HN = 67.7 mg KOH/g and basic N = 1.40%.

[00085] Example 3 (N,N-dibutylamino-N-ethyl lauryl+3 EO-ether carboxylate)
161.5 g of lauryl+3 EO-ether carboxylic acid (HN = 104.2 mg KOH/g) and 52.0 g of dibutylaminoethanol (DBAE, OHN = 323.8 mg KOH/g) were used to obtain approx. 190 g of N,N-dibutylamino-N-ethyl lauryl+3 EO-ether carboxylate having AN = 4.4 mg KOH/g, HN = 98.7 mg KOH/g and basic N = 1.91%.

[00086] Example 4 (N,N-dibutylamino-N-ethyl 2-ethylhexyl+5 EO-ether carboxylate)
188.5 g of 2-ethylhexyl+5 EO-ether carboxylic acid (HN = 89.3 mg KOH/g) and 52.0 g of dibutylaminoethanol (DBAE, OHN = 323.8 mg KOH/g) were used to obtain approx. 215 g of N,N-dibutylamino-N-ethyl 2-ethylhexyl+5 EO-ether carboxylate having AN = 6.5 mg KOH/g, HN = 93.9 mg KOH/g and basic N = 1.80%.

[00087] Example 5 (N,N-dibutylamino-N-ethyl 2-ethylhexyl+3 EO-ether carboxylate)
142.2 g of 2-ethylhexyl+3 EO-ether carboxylic acid (HN = 118.4 mg KOH/g) and 52.0 g of dibutylaminoethanol (DBAE, OHN = 323.8 mg KOH/g) were used to obtain approx. 175 g of N,N-dibutylamino-N-ethyl 2-ethylhexyl+3 EO-ether carboxylate having AN = 1.4 mg KOH/g, HN = 104.0 mg KOH/g and basic N = 2.21%.

[00088] Example 6 (N,N-dibutylamino-N-ethyl C_{14/15} oxo alcohol+7 EO-ether carboxylate)
220.1 g of C_{14/15} oxo alcohol+7 EO-ether carboxylic acid (HN = 76.5 mg KOH/g) and 52.0 g of dibutylaminoethanol (DBAE, OHN = 323.8 mg KOH/g) were used to obtain approx. 245 g of N,N-dibutylamino-N-ethyl

C_{14/15} oxo alcohol+7 EO-ether carboxylate having AN = 6.2 mg KOH/g, HN = 78.3 mg KOH/g and basic N = 1.43%.

- [00089] General method for the preparation of N,N-dialkylaminoalkoxy ether carboxylate
- [00090] A stirred apparatus equipped with a reflux condenser was initially charged with 0.3 mol of the appropriate ether carboxylic acid (based on hydrolysis number) with nitrogen purging and heated to 60°C. 0.3 mol of the appropriate N,N-dialkylaminopolyglycol and 0.5% by weight of butylstannic acid were then added. The reaction mixture was heated to 120°C, in the course of which the first water of condensation passed over. The reaction temperature was then increased continuously to 180°C over 15 h until no more water of reaction condensed out and an acid number of < 10 mg KOH/g had been attained. Finally, the reaction product was clarified using a pressure filter press with Seitz T 2600.
- [00091] Example 7 (N,N-dibutylamino-N-tri(ethoxy)ethyl oleyl+2 EO-ether carboxylate)
163.3 g of oleyl+2 EO-ether carboxylic acid (HN = 103.1 mg KOH/g) and 72.9 g of dibutylamine+3.7 EO (OHN = 230.9 mg KOH/g) were used to obtain approx. 215 g of N,N-dibutylamino-N-tri(ethoxy)ethyl oleyl+2 EO-ether carboxylate having AN = 5.1 mg KOH/g, HN = 80.3 mg KOH/g and basic N = 1.54%.
- [00092] Example 8 (N,N-dibutylamino-N-tri(ethoxy)ethyl lauryl+3 EO-ether carboxylate)
161.5 g of lauryl+3 EO-ether carboxylic acid (HN = 104.2 mg KOH/g) and 72.9 g of dibutylamine+3.7 EO (OHN = 230.9 mg KOH/g) were used to obtain approx. 205 g of N,N-dibutylamino-N-tri(ethoxy)ethyl lauryl+3 EO-ether carboxylate having AN = 6.0 mg KOH/g, HN = 81.4 mg KOH/g and basic N = 1.66%.

- [00093] Example 9 (N,N-dibutylamino-N-tri(ethoxy)ethyl 2-ethylhexyl+3 EO-ether carboxylate)
142.2 g of 2-ethylhexyl+3 EO-ether carboxylic acid (HN = 118.4 mg KOH/g) and 72.9 g of dibutylamine+3.7 EO (OHN = 230.9 mg KOH/g) were used to obtain approx. 195 g of N,N-dibutylamino-N-tri(ethoxy)ethyl 2-ethylhexyl+3 EO-ether carboxylate having AN = 4.7 mg KOH/g, HN = 87.6 mg KOH/g and basic N = 1.85%.
- [00094] Example 10 (N,N-dibutylamino-N-tri(ethoxy)ethyl C_{14/15}-oxo alcohol+7 EO-ether carboxylate)
220.1 g of C_{14/15}-oxo alcohol+7 EO-ether carboxylic acid (HN = 76.5 mg KOH/g) and 72.9 g of dibutylamine+3.7 EO (OHN = 230.9 mg KOH/g) were used to obtain approx. 265 g of N,N-dibutylamino-N-tri(ethoxy)ethyl C_{14/15}-oxo alcohol+7 EO-ether carboxylate having AN = 8.5 mg KOH/g, HN = 57.4 mg KOH/g and basic N = 1.33%.
- [00095] General method for the preparation of N,N-dialkylaminoalkyl ether carboxamides
- [00096] A stirred apparatus equipped with a reflux condenser was initially charged with 0.3 mol of the appropriate ether carboxylic acid (based on hydrolysis number) with nitrogen purging and heated to 60°C. 0.3 mol of the appropriate N,N-dialkylaminoalkylamine was then added and the reaction mixture was heated to 120°C, in the course of which the first water of condensation passed over. The reaction temperature was then increased continuously to 200°C over 5 h until no more water of reaction condensed out and an acid number of < 10 mg KOH/g had been attained. Finally, the reaction product was clarified using a pressure filter press with Seitz T 2600.
- [00097] Example 11 (N,N-dimethylamino-N-propyl oleyl+5 EO-ether

carboxamide)

203.9 g of oleyl+5 EO-ether carboxylic acid (HN = 82.6 mg KOH/g) and 30.7 g of dimethylaminopropylamine (DMAPA) were used to obtain approx. 200 g of N,N-dimethylamino-N-propyl oleyl+5 EO-ether carboxamide having AN = 6.4 mg KOH/g, amide N = 1.88% and basic N = 1.82%.

[00098] Example 12 (N,N-dimethylamino-N-propyl lauryl+3 EO-ether carboxamide)

161.5 g of lauryl+3 EO-ether carboxylic acid (HN = 104.2 mg KOH/g) and 30.7 g of dimethylaminopropylamine (DMAPA) were used to obtain approx. 180 g of N,N-dimethylamino-N-propyl lauryl+3 EO-ether carboxamide having AN = 6.9 mg KOH/g, amide N = 2.14% and basic N = 2.06%.

[00099] Example 13 (N,N-dimethylamino-N-propyl C_{14/15}-oxo alcohol+7 EO-ether carboxamide)

220.1 g of C_{14/15}-oxo alcohol+7 EO-ether carboxylic acid (HN = 76.5 mg KOH/g) and 30.7 g of dimethylaminopropylamine (DMAPA) were used to obtain approx. 275 g of N,N-dimethylamino-N-propyl C_{14/15}-oxo alcohol+7 EO-ether carboxamide having AN = 5.1 mg KOH/g, amide N = 1.51% and basic N = 1.51%.

[000100] Example 14 (N,N-dibutylamino-N-propyl C_{14/15}-oxo alcohol+7 EO-ether carboxamide)

220.1 g of C_{14/15}-oxo alcohol+7 EO-ether carboxylic acid (HN = 76.5 mg KOH/g) and 55.9 g of dibutylaminopropylamine (DBAPA) were used to obtain approx. 255 g of N,N-dibutylamino-N-propyl C_{14/15}-oxo alcohol+7 EO-ether carboxamide having AN = 3.8 mg KOH/g, amide N = 1.42% and basic N = 1.39%.

[000101] Example 15 (N,N-dibutylamino-N-propyl 2-ethylhexyl+3 EO-ether

carboxamide)

142.2 g of 2-ethylhexyl+3 EO-ether carboxylic acid (HN = 118.4 mg KOH/g) and 55.9 g of dibutylaminopropylamine (DBAPA) were used to obtain approx. 175 g of N,N-dibutylamino-N-propyl 2-ethylhexyl+3 EO-ether carboxamide having AN = 5.7 mg KOH/g, amide N = 2.00% and basic N = 1.96%.

[000102] Example 16 (N,N-dibutylamino-N-ethyl C_{14/15}-oxo alcohol+7 EO-ether carboxamide)

220.1 g of C_{14/15}-oxo alcohol+7 EO-ether carboxylic acid (HN = 76.5 mg KOH/g) and 51.7 g of dibutylaminoethylamine (DBAEA) were used to obtain approx. 255 g of N,N-dibutylamino-N-ethyl C_{14/15}-oxo alcohol+7 EO-ether carboxamide having AN = 5.4 mg KOH/g, amide N = 1.42% and basic N = 1.45%.

[000103] General method for the quaternization of dimethyl sulfate

[000104] A stirred apparatus was initially charged with 0.2 mol (based on basic N) of the appropriate tertiary amine with nitrogen purging and heated to 60°C. 0.19 mol of dimethyl sulfate was added dropwise thereto in such a way that the reaction temperature did not exceed 80-90°C. The reaction mixture was subsequently stirred at 90°C for a further 3 h. This method was used to quaternize the compounds described by examples 1 to 16 (examples 17 to 32, as listed in tables 1 to 4).

[000105] General method for the preparation of the betaines

[000106] A stirred apparatus was initially charged with 0.2 mol (based on basic N) of the appropriate tertiary amine with nitrogen purging and dissolved in 10% by weight of isopropanol (based on the total amount) at 60°C with continuous stirring. 0.25 mol of monochloroacetic acid was then added in one portion and stirred homogeneously. Subsequently, 0.27 mol of an aqueous NaOH solution (21.6 g of a 50% solution) was

allowed to flow in 4 portions into this reaction mixture. The addition was effected in such a way that the internal temperature did not exceed 60 - 65°C (in some cases, cooling was necessary). After each addition, reaction was continued at 80°C in each case for 30 minutes, and finally for 3 hours. Finally, the precipitated NaCl residue was removed by means of a pressure filter press through Seitz T 5500.

[000107] Example 33 (oleyl+5 EO-ether carboxamido-N,N-dimethylamino-N-propylammonium N-methylcarboxyl betaine)

153.8 g of N,N-dimethylamino-N-propyl oleyl+5 EO-ether carboxamide (basic N = 1.82%) and 23.6 g of monochloroacetic acid (MCAA) were used to obtain approx. 150 g of oleyl+5 EO-ether carboxamido-N,N-dimethylamino-N-propylammonium N-methylcarboxyl betaine (AS content approx. 83%, approx. 7% water, approx. 10% isopropanol).

[000108] Example 34 (C_{14/15}-oxo alcohol+7 EO-ether carboxamido-N,N-dimethylamino-N-propylammonium N-methylcarboxyl betaine)

185.4 g of N,N-dimethylamino-N-propyl C_{14/15}-oxo alcohol+5 EO-ether carboxamide- (basic N = 1.51%) and 23.6 g of monochloroacetic acid (MCAA) were used to obtain approx. 195 g of C_{14/15}-oxo alcohol+5 EO-ether carboxamido-N,N-dimethylamino-N-propylammonium N-methylcarboxyl betaine (AS content approx. 84%, approx. 6% water, approx. 10% isopropanol).

[000109] Example 35 (lauryl+3 EO-ether carboxamido-N,N-dimethylamino-N-propylammonium N-methylcarboxyl betaine)

135.9 g of N,N-dimethylamino-N-propyl lauryl+3 EO-ether carboxamide- (basic N = 2.06%) and 23.6 g of monochloroacetic acid (MCAA) were used to obtain approx. 130 g of lauryl+3 EO-ether carboxamido-N,N-dimethylamino-N-propylammonium N-methylcarboxyl betaine (AS content approx. 82%, approx. 8% water, approx. 10% isopropanol).

Example 36

- [000110] Polyvinylcaprolactam having MW 5 000 g/mol is mixed in a ratio of 1:1 with the quats of example 9 described by example 25 and terminated in butyldiglycol.

Example 37

- [000111] Polyvinylcaprolactam having MW 5 000 g/mol is mixed in a ratio of 1:1 with the quats of example 15 described by example 31 and terminated in butyldiglycol.
- [000112] Effectiveness of the compounds according to the invention as corrosion inhibitors
- [000113] The compounds according to the invention were tested as corrosion inhibitors in the Shell wheel test. Coupons of carbon steel (DIN 1.1203 having 15 cm² surface area) were immersed in a salt water/petroleum mixture (9:1.5% NaCl solution, adjusted to pH 3.5 using acetic acid) and subjected to this medium at a rotation rate of 40 rpm at 70°C for 24 hours. The dosage of the inhibitor was 50 ppm of a 40% solution of the inhibitor. The protection values were calculated from the mass reduction of the coupons, based on a blank value.
- [000114] In the table which follows, "comparison 1" refers to a commercial residue amine, quat based on dicocoalkyldimethylammonium chloride and "comparison 2" to a commercial soya fatty acid amidopropyl-N,N-dimethylammonium carboxymethyl betaine described by EP-B-0 446 616 (prior art corrosion inhibitors).

Table 1: (SHELL wheel test)

Example	Corrosion inhibitor	ø Protection %
Comparison 1	Standard quat	36.0
Comparison 2	Standard betaine	75.4
17	Quat from example 1	91.2
18	Quat from example 2	93.2
19	Quat from example 3	92.8
22	Quat from example 6	84.1
23	Quat from example 7	93.7
24	Quat from example 8	93.0
26	Quat from example 10	88.5
27	Quat from example 11	78.6
28	Quat from example 12	70.5
29	Quat from example 13	73.6
30	Quat from example 14	94.2
32	Quat from example 16	92.9
33	Betaine from example 11	70.7
34	Betaine from example 12	71.9
35	Betaine from example 13	75.9

The products were also tested in the LPR test (test conditions similar to ASTM D 2776).

Table 2: (LPR test)

Example	Corrosion inhibitor	Protection after [%]		
		10 min	30 min	60 min
Comparison 1	Standard quat	53.9	61.2	73.7
Comparison 2	Standard betaine	45.9	59.2	64.3
17	Quat from example 1	76.4	85.8	90.2
18	Quat from example 2	69.6	77.0	81.4
19	Quat from example 3	85.4	91.5	93.1
22	Quat from example 6	82.3	88.1	89.5
23	Quat from example 7	86.8	92.7	94.1
26	Quat from example 10	78.0	86.9	91.3
33	Betaine from example 11	75.6	80.3	82.8
35	Betaine from example 13	74.4	80.2	84.3

[000115] As can be seen from the above test results, the products according to the invention have very good corrosion protection properties at low dosage and substantially exceed the effectiveness of the prior art inhibitors. As a consequence of their ester and amide structure, the compounds have better biodegradability and can be used at a lower dosage.

[000116] Effectiveness of the compounds according to the invention as gas hydrate inhibitors

[000117] To investigate the inhibiting action of the compounds according to the invention, a stirred steel autoclave having temperature control, pressure and torque sensor having an internal volume of 450 ml was used. For investigations of kinetic inhibition, the autoclave was filled with distilled water and gas in a volume ratio of 20:80, and for investigations of agglomerate inhibition, condensate was additionally added. Finally, natural gas was injected at different pressures.

[000118] Starting from a starting temperature of 17.5°C, the autoclave was

cooled to 2°C within 2 h, then stirred at 2°C for 18 and heated back up to 17.5°C within 2 h. A pressure decrease corresponding to the thermal compression of the gas is observed. When the formation of gas hydrate nuclei occurs during the supercooling time, the pressure measured falls, and an increase in the torque measured and a slight increase in temperature can be observed. Without inhibitor, further growth and increasing agglomeration of the hydrate nuclei lead rapidly to a further increase in the torque. When the mixture is heated, the gas hydrates decompose, so that the starting state of the experimental series is attained.

[000119] The measure used for the inhibiting action of the compounds according to the invention is the time from the attainment of the minimum temperature of 2°C up to the first gas absorption (T_{ind}) or the time up to the rise of the torque (T_{agg}). Long induction times or agglomeration times indicate action as a kinetic inhibitor. On the other hand, the torque measured in the autoclave serves as a parameter for the agglomeration of the hydrate crystals. The pressure drop measured (Δp) allows a direct conclusion on the amount of hydrate crystals formed. In the case of a good antiagglomerate, the torque which builds up after formation of gas hydrates is distinctly reduced compared to the blank value. Ideally, the snowlike, fine hydrate crystals form in the condensate phase and do not agglomerate and thus lead to blockage of the installations serving for gas transport and for gas extraction.

Test results

Composition of the natural gas used:

[000120] Gas 1: 79.3% methane, 10.8% ethane, 4.8% propane, 1.9% butane, 1.4% carbon dioxide, 1.8% nitrogen. Supercooling below the equilibrium temperature of hydrate formation at 50 bar: 12°C.

[000121] Gas 2: 92.1% methane, 3.5% ethane, 0.8% propane, 0.7% butane, 0.6% carbon dioxide, 2.3% nitrogen. Supercooling below the equilibrium temperature of hydrate formation at 50 bar: 7°C, supercooling at 100 bar: 12°C.

[000122] In order to test the effectiveness as agglomerate inhibitors, the test autoclave used above was initially charged with water and white spirit (20% of the volume in a ratio of 1:2) and, based on the aqueous phase, 5 000 ppm of the particular additive were added. At an autoclave pressure of 90 bar using gas 1 and a stirred speed of 5 000 rpm, the temperature was cooled from initially 17.5°C within 2 hours to 2°C, then the mixture was stirred at 2°C for 25 hours and heated again. The pressure drop caused by hydrate formation and the resulting torque at the stirrer, which is a measure of the agglomeration of the gas hydrates, were measured.

Table 3: (Test as antiagglomerant)

Example	Gas hydrate inhibitor	Pressure drop Δp (bar)	Temperature rise ΔT (K)	Torque M_{\max} (Ncm)
Blank value	-	> 40	> 8	15.9
19	Quat from example 3	15.8	0.8	0.5
20	Quat from example 4	15.3	1.0	6.4
21	Quat from example 5	10.8	0.1	0.2
25	Quat from example 9	9.1	0.0	0.0
31	Quat from example 15	10.4	0.2	0.3
Comparison 3	TBAB	21.5	1.0	1.5
Comparison 4	TBAB	15.0	1.0	1.2

[000123] The comparison substances 3 and 4 used were two commercially available antiagglomerant inhibitors based on tetrabutylammonium

bromide (TBAB).

[000124] As can be seen from these examples, the torques measured were greatly reduced in comparison to the blank value despite severe hydrate formation. This supports a distinct agglomerate-inhibiting action of the products according to the invention. It is apparent that especially in the case of balanced HL balance, excellent results are achieved.

[000125] In order to test the effectiveness as additives for kinetic inhibitors, 5 000 ppm of the particular additive, based on the aqueous phase, were added in the above-described test autoclave and cooled at different pressures using gases 1 or 2. On attainment of the minimum temperature of 2°C, the time until the first gas absorption (T_{ind}) was recorded. The pressure drop (Δp) and the temperature rise ΔT (K) allow the amount of hydrate crystals formed to be concluded directly.

Table 4: (Test as kinetic inhibitors)

Example	Inhibitor	Gas	Pressure p (bar)	T_{ind}	Pressure drop Δp (bar)	Temperature rise ΔT (K)
Blank value	-	1	50	0	> 40	> 1.5
Blank value	-	2	100	0	> 40	> 1.5
36	Quat from example 9	1	50	11.3 h	9.1	0.2
36	Quat from example 9	2	100	< 5 min	10.8	0.4
37	Quat from example 15	1	50	17.0 h	3.5	0.1
37	Quat from example 15	2	100	1.5 h	5.5	0.3
Comparison 5	PVCap	1	50	< 5 min	10	0.4
Comparison 5	PVCap	2	100	< 5 min	6	0.1

[000126] The comparison substance 5 used was a solution of polyvinylcaprolactam (PVCap) in butylglycol, molecular weight 5 000 g/mol.

[000127] As can be recognized from the above test results, the products according to the invention act as a synergistic component of kinetic hydrate inhibitors and exhibit a distinct improvement compared to the prior art. They can therefore be used for increasing (synergistic effect) the performance of prior art inhibitors.

The following abbreviations have been used:

AN = acid value (alternatively "acid number")

HN = saponification value (alternatively "saponification number")

OHN = hydroxyl value (alternatively "hydroxyl number")

AS content = content of active matter